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 $E_{cell} = E_{red}(Cu^{2+}) - E_{ox}(Ag) = 0.34V - 0.80V = -0.46V$ . Not spontaneous. ox:  
 $[Ni(s) \rightarrow Ni^{2+} + 2e^-] \times 5$ ; red:  $[5e^- + 8H^+ + MnO_4^- \rightarrow Mn^{2+} + 4H_2O] \times 2$ .



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